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1. Similarities: Both types of bonds are the result of overlapping atomic orbitals on adjacent atoms and contain a maximum of two electrons. Differences:  $\sigma$  bonds are stronger and are the result of overlap, and all individual bonds  $\sigma$  bonds;  $\pi$  between the same two atoms are weaker because they are the result of overlapping side by side, and multiple bonds contain one or more  $\pi$  bonds (in addition to  $\sigma$  bond). 3. A specific average distance of communication is the distance with the least energy. At distances shorter than the distance of communication, positive charges on the two nuclei repel each other, and the total energy increases. 5. Communication: one  $\sigma$  and one  $\pi$  connection. Orbitals are filled and do not intersect. Orbitals overlap along the axis to form a  $\sigma$  connection and side by side to form a  $\pi$  connection. 7. No, two p orbitals (one on each N) will be end-to-end and form  $\sigma$  connection. 9. Hybridization is introduced to explain the geometry of the bonding of orbital stations in the theory of valence bonds. There are no d orbiting stations in the carbon valence shell. 13. trigonal planar, sp<sup>2</sup>; trigonal pyramidal (one lone pair on A) sp<sup>3</sup>; T-shaped (two single pairs on A) sp<sup>3</sup>d, or (three single pairs on A) sp<sup>3</sup>d<sup>2</sup> 15. a) each S has bent (109°) geometry, sp<sup>3</sup> (b) Bent (120°), sp<sup>2</sup> (c) Trigonal planar, sp<sup>2</sup> (d) Tetrahedral, sp<sup>3</sup> 17. (a) XeF<sub>2</sub> (b) (c) linear (d) sp<sup>3</sup>d 19. (a) P atoms, trigonal pyramidal; S atoms, bent, with two lone pairs; Cl atoms, trigonal pyramidal; (c) Hybridization of P, S and Cl is in all cases sp<sup>3</sup>; (d) Oxidation reads P No.1, S-113, No113, Cl No5, O-2. Official charges: P 0; S 0; Cl No2; O -1 21. Phosphorus and nitrogen can form sp<sup>3</sup> hybrids to form three bonds and hold one single pair in PF<sub>3</sub> and NF<sub>3</sub>, respectively. However, nitrogen has no valence d orbital, so it cannot form a set of sp<sup>3</sup>d hybrid orbitals to bind five fluoride atoms in NF<sub>5</sub>. Phosphorus has a d orbital and can bind five fluoride atoms with sp<sup>3</sup>d hybrid orbital in PF<sub>5</sub>. The triple bond consists of one  $\sigma$  bond and two  $\pi$  bonds. More  $\sigma$  bonds are stronger than  $\pi$  due to more duplication. 25. (a) (b) The terminal carbon atom uses hybrid sp<sup>3</sup> orbital systems, while the central carbon atom is hybridized. (c) Each of the two  $\pi$  of communications is formed by the overlap of the 2p orbit on carbon and nitrogen orbit 2p.27. (a) sp<sup>2</sup>; b) sp; c) sp<sup>2</sup>; d) sp<sup>3</sup>; e) sp<sup>3</sup>; f) sp<sup>3</sup>d; g) sp<sup>3</sup> 29. (a) sp<sup>2</sup>, delocalized; (b) sp, localized; (c) sp<sup>2</sup>, delocalized; (d) sp<sup>3</sup>, delocalized 31. Each of the four electrons is in a separate orbit and is covered by an electron on the oxygen atom 33. (a) Similarities: Both are bound by orbital orbitals that can contain a maximum of two electrons. Differences:  $\sigma$  orbitals are endless combinations of atomic orbital stations, while  $\pi$  orbitals are formed side by side by the overlapping orbital stations. Similarities: Both structures represent the probability of finding an electron about an atom or molecule. Differences: the  $\psi$  for atomic orbit describes the behavior of only one electron at a time, time, Atom. For the molecule,  $\psi$  is a mathematical combination of atomic orbiting stations. (c) Similarities: Both are orbitals that may contain two electrons. Differences: Gluing orbitals results in two or more atoms together. Anti-bond orbital systems can destabilize any bonding that has occurred. 35. An odd number of electrons can never be paired, regardless of the location of molecular orbitals. It will always be paramagnetic. 37. Orbital communications have a density of electrons in close proximity to more than one nucleus. The interaction between the bonding of positively charged nuclei and negatively charged electrons stabilizes the system. 39. Pairing two bonding electrons reduces the energy of the system in relation to the energy of non-renewable electrons. 41. (a) order of H<sub>2</sub> Bonds No. 1, H<sub>2</sub>H<sub>2</sub> bond order - 0.5, H<sub>2</sub>-H<sub>2</sub> bonds - 0.5, strongest bond - H<sub>2</sub>; (b) O<sub>2</sub> No 2 bond order, O<sub>2</sub>O<sub>2</sub> 3 bond order, Order for O<sub>2</sub>-O<sub>2</sub> bonds - 1, strongest bond - O<sub>2</sub>O<sub>2</sub>; (c) Order for Bonds Li<sub>2</sub> No 1, Be<sub>2</sub>Be<sub>2</sub> Bond Order - 0.5, Be<sub>2</sub> Bond Order - 0, Strongest Bond - Li<sub>2</sub>Li<sub>2</sub>; (d) F<sub>2</sub> Bond Order - 1, F<sub>2</sub>-F<sub>2</sub> - Bond Order - 1.5, F<sub>2</sub>-F<sub>2</sub> - 0.5, Strongest Bond - F<sub>2</sub>; F<sub>2</sub>; (e) N<sub>2</sub> bonds of about No 3, N<sub>2</sub> 'N<sub>2</sub> bonds about 2.5, N<sub>2</sub>N<sub>2</sub> bonds about 2.5, strongest bonds N<sub>2</sub>43. (a) H<sub>2</sub>; (b) N<sub>2</sub>; (c) O<sub>2</sub>; (d) C<sub>2</sub>; (e) B<sub>2</sub> 45. Yes, fluoride is smaller than Lee's, so the atoms in orbit 2s are closer to the nucleus and more stable. 49. N<sub>2</sub> has s-p mixing, so  $\pi$  orbitals are the last filled in N<sub>2</sub>. N<sub>2</sub>. The O<sub>2</sub> has no s-p mixing, so the P orbit is filled before  $\pi$  orbit. Chapter 8 Chemistry Class 12 Exercise Solutions come in an affordable version of the PDF for free download. Versions of the PDF version of CBSE NCERT Solutions are readily available on various e-learning websites such as Vedantu. They are highly effective for students in short preparation time as they can easily revise through important issues and decisions presented in detail. Decisions are also decided by experts. NCERT Solution d And F Block Class 12 Examples For student convenience, solutions are divided into two sections consisting of NCERT questions and text questions. These two sections cover all the most important areas of this chapter. Here are some examples of the Class 12 d f element: In both of these questions, solutions provide a detailed explanation of the electronic configurations of the connections. It also details the reason for the stability of some of the other connections on the examples. Here, solutions provide a deep understanding of how the electronic configuration solves the stability of the state of oxidation in the first series of the transitional element. By reading these decisions, students will be able to gain control over the basic concepts in the chapters. Similarly, sections on the issue in the text also provide a detailed understanding of concepts based on examples and diagrams. Ch 8 Chemistry Class 12 SolutionsThe solution to issue 1 of Chapter 8 Chemistry Class 12 explains how a silver atom with filled d orbitals in its terrestrial state is a transitional element. The answer is explained by an adequate example of finding out all the doubts of students. Addressing this issue chapter 8 Chemistry Class 12 will help the student understand why in the Sc- (No. 21) series in enthalpy zinc spraying is the lowest. The solution uses simple language and therefore good quick understanding by students. Issue three explains to students the 3D series of transient metals that show the greatest amount of oxidation state and why. By studying this question, the student will easily answer short questions and increase his score. The solution of element d f block 12 for question 4 explains with a detailed example of why the value of E(M<sup>2+</sup> / M) for copper is positive (0.34V). The solution is given a step-by-step format for the convenience of students. By studying this method, the student will be able to remember the solution easily. (Image to be uploaded soon) Signs Distribution for Chapter 8 Chemistry Class 12B exam unit vi, vii and viii goes under 1 section and carries a total of 19 marks. 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In this chapter, students learn the concepts of the classic idea of redox reactions - Oxidation and Reaction Reduction, Redox Reactions in terms of Electron Transmission Reactions, Competitive Electron Transmission Reactions, Oxidation Room, Redox Reactions Balancing, Redox Reactions as The Basis for Titration, Restrictions on The Oxidation Concept Room, Redox Reactions and Electrode reactions. NCERT Score 11 Chemistry Part 2 Chapter 8, Redox Reactions is part of Group 8, Group 8, Group 9, Group 10 and Group 11 have a total weight of 16 marks on the final exam. Get a 100 percent accurate NCERT solution for The 8th Grade 8 Chemical Chapter (Redox Reactions) decided by chemistry experts. We provide solutions for questions asked in the Class 11 chemistry textbook in accordance with the CBSE Board guidelines from NCERT's latest book for Grade 11 Chemistry. 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